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PLAY. Hydrocarbons.

the simplest organic

compounds contain

only carbon and

hydrogen. Alkane. a

hydrocarbon in which

there are only single

covalent bonds.

Straight-chain Alkanes.

contain any number of

carbon atoms, one

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after the other.
Homologous Series.

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This hydrocarbon is called "2,3-pentadiene" because there are 2 (di) double bonds (ene) after 2 and 3 (2,3) of a 5 carbon line (penta). It is named the same way for triple bonds, but with an yne instead of ene. Next are isomers. Isomers are

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compounds with the
same molecular
formula but different
molecular structures.

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Corcoran. How is the

number of valence

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electrons in carbon atoms related to the bonds that carbon atoms form? because carbon has four valence electrons, a carbon atom always forms four covalent bonds.

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these activities to learn the vocabulary and major concepts presented in the chapter. molecules in which the atoms are joined in the same order, but the arrangements of the atoms in space are different.

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22.1 Hydrocarbons. •
Because carbon has
four valence electrons,
car- bon atoms always
form four covalent
bonds. • The carbon
atoms in an alkane can
be arranged in a
straight chain or in a
chain that has
branches. • Molecules

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of hydrocarbons, such as alkanes, are nonpolar molecules.

22.2 Unsaturated Hydrocarbons.

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Hydrocarbons and
Their Reactions. 784
Formic acid.

BIGIdeaThe
substitution of different
functional groups for
hydrogen atoms in
hydrocarbons results in
a diverse group of
organic compounds.

22.1 Alkyl Halides and
Aryl Halides.

MAINIdeaA halogen
atom can replace a
hydrogen atom in

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some hydrocarbons.

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introduction to
functional groups.

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Compounds. kerosene.
Diesel fuel. gasoline. 2.
Organic Chemistry
Hydrocarbons.
Scientists used to think
that only living.

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organisms could synthesize the carbon compounds. found in their cells.

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Unsaturated

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